

Article

# Tailored Oxidation Barrier Coatings for Mo-Hf-B and Mo-Zr-B Alloys

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**Abstract:** The cyclic oxidation response of Mo-14Hf-23B and Mo-14.8Zr-26B (compositions in at. %) was investigated in air at 800 °C, which is a critical temperature for Mo-based alloys because of the pesting phenomenon. Rapid oxidation was observed for the unprotected samples, and an oxidation protection coating was developed based on a preceramic polymer with silicon and boron as particulate fillers. Cyclic oxidation tests of the coated samples showed excellent oxidation protection: no Mo, Hf or Zr oxides were found after testing and a small mass gain in the initial stage of oxidation indicated the formation of a glassy protection layer on the alloys surfaces after exposure to air at 800 °C.

**Keywords:** polymer-derived ceramics; preceramic polymers; perhydropolysilazane; oxidation barrier coatings; Mo-Hf-B; Mo-Zr-B

## 1. Introduction

One of the challenging tasks in materials science is the development of novel high-temperature materials; characteristics of high temperature structural materials, besides high melting temperatures, are a high creep resistance, an appropriate tolerance for crack initiation and crack growth, resistance against oxidative, corrosive and erosive attacks, and thermal shocks [1,2]. The properties of known high-temperature structural materials can limit the advance of machine design in airplane engines or power plants [3,4]. For example, the most common materials in high-pressure turbines of aero-engines and gas turbines—nickel-based super alloys—have reached their technological limit, and are not capable of having structural components at temperatures above 1150 °C [5]. High melting temperature metals such as molybdenum alloys represent an alternative, but they suffer from rapid oxidation and a high creep rate at high temperatures [3]. The reinforcement of the refractory metals with a stronger phase, which may preferably be oxidation resistant, might solve this problem [4]. In this context, refractory metal borides, carbides and silicides are the most appropriate candidates.

Potential examples of such materials are the novel Mo-Hf-B and Mo-Zr-B alloys [6,7]. These alloys showed excellent creep properties; their creep rates are several orders of magnitude lower compared to that of a nickel based single-crystal CMSX-10 alloy [8]. Mo-Hf-B and Mo-Zr-B alloys provide also high microhardness (>1000 HV0.3) and high fracture toughness (up to 17 MPa m<sup>0.5</sup>) [6,9], and are potential structural materials for ultra-high temperature applications. Nevertheless, the resistance to oxidation media in a wide temperature range is a decisive factor for these alloy's technical applications.

Polymer-derived ceramics (PDCs) occur as a new class of highly oxidation resistive materials for individual parts or as composite coatings [10,11]. Polysiloxanes, polycarbosilanes or polysilazanes are silicon containing precursors, which can be converted into polymer-derived ceramic coatings on steel to provide thermal shock resistance, creep resistance and thermal stability [12,13]. The use of fillers reduces the shrinkage of precursor coatings during ceramics formation, which leads to an in-crease of

the coating thickness and prevention of cracks [14,15]. There is limited data in the literature of polymer ceramic coatings on molybdenum substrate.

In this study, the oxidation properties of the Mo-14Hf-23B and Mo-14.8Zr-26B alloys were tested at 800 °C in air, thus, taking a temperature from the peaking regime into account. Besides, protective coatings made of a preceramic polymer and oxidation resistant filler particles were developed and applied to these alloys aimed to improve their oxidation resistance under these conditions. It should be noted that the obtained results show the first results of oxidation properties, and do not include the oxidation kinetics yet.

## 2. Materials and Methods

The materials investigated in this paper were manufactured by an arc melting process. The arc melted Mo-14Hf-23B and Mo-14.8Zr-26B (in at. %) alloys were produced from elemental flakes of Mo, Hf, Zr (all from E. WAGENER GmbH, Heimsheim, Germany) and B (HMW Hauner GmbH, Roettenbach, Germany) with a purity of 99.95%, 98.0%, 97.2% and 99.0%, respectively. Processing was carried out in a Buehler arc furnace of the type MAM 1 (Edmund Buehler GmbH, Bodelshausen, Germany) in high-purity argon atmosphere. The alloys were subsequently cooled in a copper mould. The melting process was carried out at least for five times including flipping of the samples in-between each melting step to improve the chemical homogeneity.

For the cyclic oxidation tests, the Mo-Hf-B and Mo-Zr-B alloys were cut by electrical discharge machining (EDM) into samples of 2 mm × 1.5 mm × 8 mm. Cyclic oxidation tests were performed in a convection furnace (Model L9/S17, Nabetherm, Lilienthal, Germany) at 800 °C. Interval withdrawal of samples for characterization was carried out after 10 min during the initial oxidation stage (<1 h), and in 1-h steps up to 20 h followed by 5-h steps up to 50 h, respectively. The specific mass change of the samples was measured using an analytical balance with an accuracy of 0.0001 g.

For the application of a preceramic polymer coating, a dip-coating method was carried out in a self-constructed dip coater; a withdrawal speed of 3 mm/s and a sample holding time of 10 s were applied. Prior to coating, the alloy samples were ground with SiC paper FEPA P1200 (17.3 µm to 19.3 µm), and cleaned in an acetone (80 vol.)/ethanol (20 vol.) mixture by ultrasonic treatment, and subsequently dried. As a preceramic polymer, a commercially available perhydropolysilazane (PHPS NN 120-20, Clariant Advanced Materials GmbH, Sulzbach, Germany,) was chosen. Silicon (Si, purity: 99.50%; Thermo Fisher GmbH, Kandel, Germany) and boron (B, purity: 95.00–97.00%; Chempur, Feinchemikalien und Forschungsbedarf GmbH, Karlsruhe, Germany) filler powders (~5 µm) were introduced into the precursor solution (20 wt.% of PHPS in dibutyl ether) in a ratio of 64 PHPS + 24 Si + 12 B, all numbers in vol.%, by vigorous stirring under air.

The as-coated samples were heated to 110 °C with a slope of 2 K/min in air for 1–1.5 h to allow crosslinking. Pyrolysis was conducted in nitrogen atmosphere in a high temperature furnace (HTRRH 70-600/18, Fa. Gero Hochtemperaturöfen GmbH & Co. KG, Neuhausen, Switzerland) for 1 h at 1000 °C peak temperature; heating and cooling rates were 3 K/min. The thickness of the coating was achieved by four times overcasting the first layer followed by crosslinking and pyrolysis after each step.

The thermal transformation with respect to weight change was analyzed with a simultaneous thermal analyzer (STA 449 F3 Jupiter, Netzsch-Gerätebau GmbH, Selb, Germany; nitrogen, 50 mL/min; heating rate was 5 K/min). The most probable chemical reactions between Mo-Zr-B alloys' constituents and perhydropolysilazane Si-N<sub>2</sub>-H<sub>2</sub> (O<sub>2</sub> in case of moisture) were formulated. Thermodynamic calculations of the free Gibb's enthalpy of each of these chemical reactions were carried out with the software package HSC Chemistry 7.16 (Outotec Oy, Pori, Finland). The equilibrium of the most expected phases in a closed system with maximum entropy was determined in the temperature range between 600 °C and 1600 °C using the same software package. The output data are presented in Table 1. The phase fractions as a function of the temperature and the pyrolysis atmosphere were calculated.

**Table 1.** Chemical composition of the system polysilazane (PHPS NN 120-20; Si–N<sub>2</sub>–H<sub>2</sub>) alloy (Mo, Zr, B) and pyrolysis atmosphere (O<sub>2</sub>, N<sub>2</sub>) used for thermodynamic calculations.

Components Amount, mol:		
Polysilazane (PHPS NN 120-20)	From the Alloy's Surface	From the Atmosphere
Ar		1.0
N <sub>2</sub>	0.596	0.200–0.400
H <sub>2</sub>	0.193	
Si	0.238	
Mo	0.082	
Zr	0.019	
B	0.036	
O <sub>2</sub>	0.001	0.010–0.100

Phase analysis was carried out on compact samples with an X-ray diffractometer (X'PERT PRO PANalytical, Malvern Panalytical GmbH, Kassel, Germany) using Co-K $\alpha$  radiation. The phase identification was carried out with the analysis software X'Pert HighScore Plus (PANalytical, Kassel, Germany). Fourier transform infrared (FT-IR) spectra of the crosslinked and pyrolyzed samples were recorded on a Bruker Optics Vertex 70v from 4000–400 cm<sup>-1</sup>, resolution of 4 cm<sup>-1</sup>, using a Platinum attenuated reflectance technique (ATR) unit.

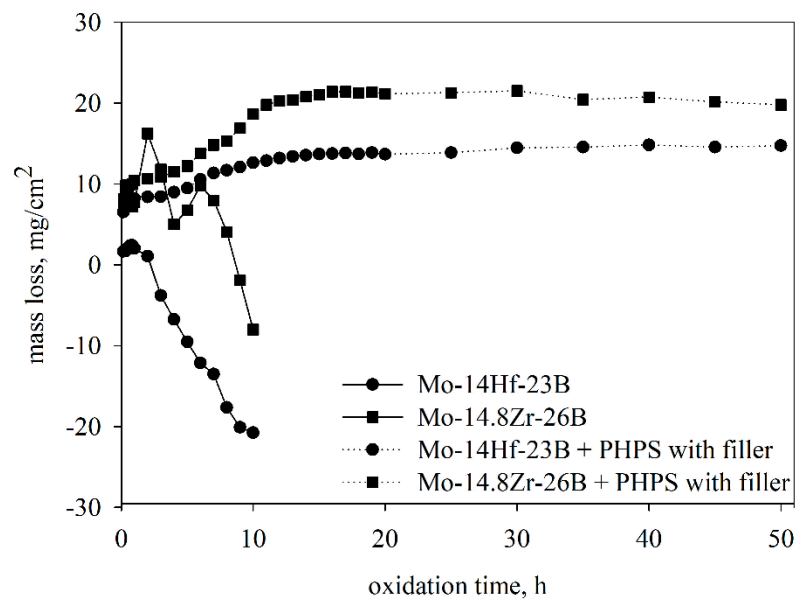
For microstructural studies, the samples were embedded in a cold-curable epoxy resin. After manual grinding from 320 down to 1200 grit size, the specimens were finished by mechanical polishing consecutively with a 3- $\mu$ m and 1- $\mu$ m diamond suspension. Scanning electron microscopy investigations (HR-FESEM, Zeiss Merlin equipped with an EDS (energy dispersive X-ray spectroscopy) detector, Carl Zeiss Microscopy GmbH, Jena, Germany; and FEI DualBeam Scios, FEI Company, Hillsboro, NH, USA) were carried out with respect to microstructure characterization and element analysis.

### 3. Results and Discussion

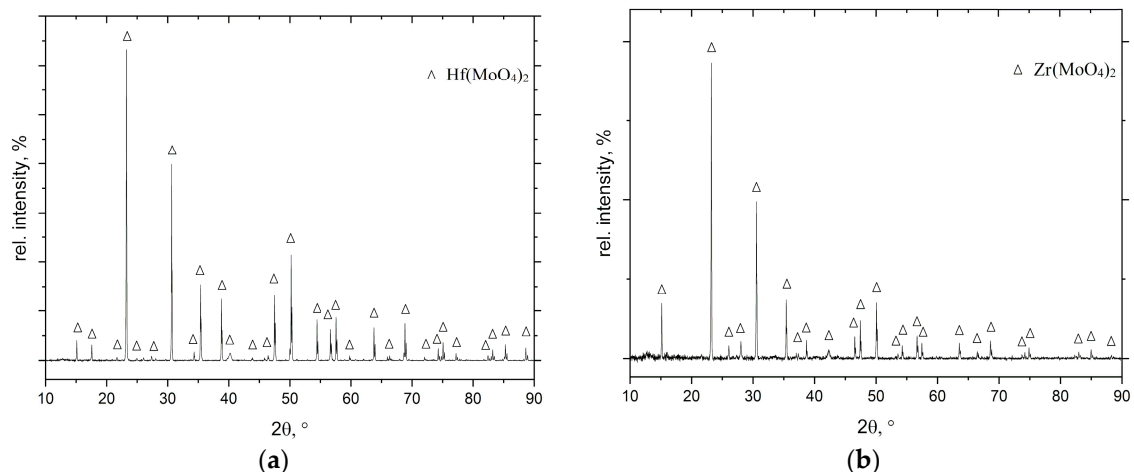
#### 3.1. Cyclic Oxidation Behavior of the Mo-14Hf-23B and Mo-14.8Zr-26B Alloys at 800 °C

The oxidation behaviour of the Mo-14Hf-23B and Mo-14.8Zr-26B alloys was cyclically proven at 800 °C. This temperature was chosen since it is the harshest temperature for Mo alloys because of the occurrence of the pesting phenomenon, which means the formation of a volatile MoO<sub>3</sub> phase [16]. During the tests, the samples failed catastrophically already within the first 10 h of exposure. The respective mass change curves are presented in Figure 1; mass change curves of coated samples are included for comparison and will be discussed later.

At the first stage (during the first 2 h), a mass gain in both alloys was observed. This oxidation step corresponds to the formation of MoO<sub>3</sub> and MoO<sub>2</sub> oxides [17]. A further oxygen penetration into the samples and evaporation of MoO<sub>3</sub> led to a rapid oxidation of the entire bulk of the alloys. The mass loss after 10 h of exposure in air is –21.6 mg/cm<sup>2</sup> for the Mo-14Hf-23B and –9.8 mg/cm<sup>2</sup> for the Mo-14.8Zr-26B alloy. After 10 h, the samples were broken due to the formation of cracks in the oxidation products. XRD (X-ray diffraction) analyses applied to the oxide surface showed the existence of Hf(MoO<sub>4</sub>)<sub>2</sub> and Zr(MoO<sub>4</sub>)<sub>2</sub> ternary oxides (Figure 2). A boron-containing oxide was not found by XRD investigations, which might be due to a possible amorphous structure of such oxides.



**Figure 1.** Mass change curves of uncoated and of coated Mo-14Hf-23B and Mo-14.8Zr-26B alloys under oxidation at 800 °C in air.

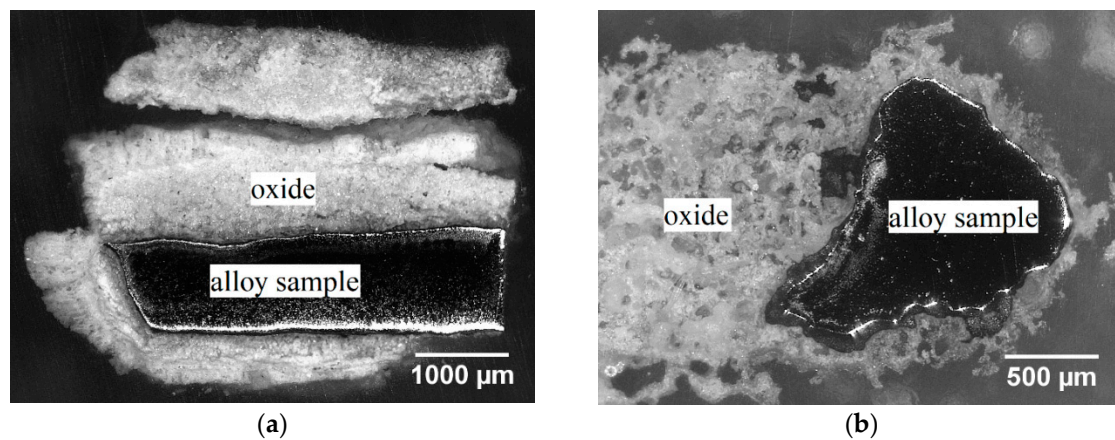


**Figure 2.** X-ray diffraction patterns of (a) Mo-14Hf-23B and (b) Mo-14.8Zr-26B alloys, both after exposure to air for 10 h at 800 °C.

After the oxidation experiment, only a few islands of the substrate were found in the cross-sections of the samples, as shown in Figure 3.

Summarizing these first oxidation experiments, both Mo-14Hf-23B and Mo-14.8Zr-26B alloys have a poor oxidation resistance at 800 °C. This points out that the use of these alloys for high temperature applications is limited to an inert or protective atmosphere. A similar oxidation behaviour was found for other Mo-based alloys [16–18]: A high volume fraction of the Mo solid solution phase in Mo-Si-B alloys affects the MoO<sub>3</sub> formation and its evaporation in a temperature range from 400 °C to 800 °C [17]. There is an exception for high Si-containing Mo-Si-B alloys, which form a SiO<sub>2</sub>-B<sub>2</sub>O<sub>3</sub> passive glassy layer by their constituents as a result of oxygen incorporation at high temperatures [16]. A major drawback of these material concepts is that in the very first heating procedure the SiO<sub>2</sub> formation is expected at temperatures above 650 °C, when the MoO<sub>3</sub> evaporation has already started, and the viscosity of the glassy layer is too high to fill gaps and holes in the materials surface after MoO<sub>3</sub>

volatilization [18]. Therefore, the oxidation of a certain amount of material from the surface cannot be avoided.



**Figure 3.** Optical microscope images of the cross-section of (a) Mo-14Hf-23B and (b) Mo-14.8Zr-26B alloys after cyclic oxidation in air at 800 °C for 10 h.

### 3.2. Oxidation Barrier Coatings

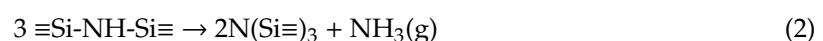
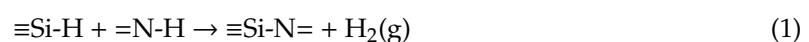
#### 3.2.1. Thermal Analysis

A novel concept of oxidation protection of Mo-Si-B based alloys, besides other concepts [18–21], is the use of filler-loaded polymeric Si precursors and their thermal conversion into ceramic layers. In [22,23], the polymeric precursor (also addressed as preceramic polymer) perhydropolysilazane (PHPS) was part of a coating system that demonstrated high temperature oxidation resistance at 800 °C and 1100 °C up to 100 h of exposure in air. After an initial thermal treatment in N<sub>2</sub> atmosphere, the PHPS coating contains mainly Si<sub>3</sub>N<sub>4</sub> or SiO (N) amorphous ceramics and elemental Si; the latter is able to trap oxygen to form a SiO<sub>2</sub> passivation layer. Passivation layer formation provides a parabolic oxidation kinetics [10,12,13]. This oxidation protection mechanism, based on a thermally stable SiO<sub>2</sub> layer on the outside, might become a profitable alternative for the improvement of oxidation resistance of Mo-Hf-B and Mo-Zr-B alloys.

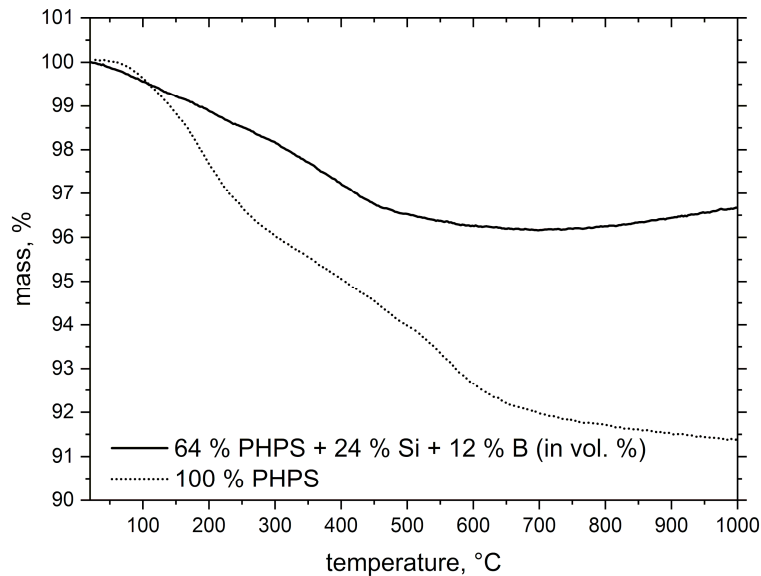
Following the idea to provide sufficient B<sub>2</sub>O<sub>3</sub> to the SiO<sub>2</sub> passive glassy layer for viscosity adjustment [16], Si and B particulate fillers were added to the preceramic polymer coating system; a glassy SiO<sub>2</sub>-B<sub>2</sub>O<sub>3</sub> layer may have a good adherence to the substrate material, and may serve to overcome the mismatch of the coefficients of thermal expansion between coating and alloys.

For the coating system, a slurry composition of 64 PHPS + 24 Si + 12 B, all numbers in vol.%, was chosen. The resulting Si/B ratio was predetermined by experimental data [22,23], where the above-mentioned type of coating system was applied to Mo-Si-B alloys.

The thermogravimetric analysis of the coating slurry is represented in Figure 4. The thermal transformation curves demonstrate significantly reduced mass loss of the particle-filled PHPS in comparison to the plain PHPS. The main mass loss of −8.6 wt.% of the previously crosslinked PHPS was observed in the temperature range up to 700 °C—this is due to the release of H<sub>2</sub> and NH<sub>3</sub>, Equations (1) and (2), during thermally induced structural rearrangement, and the conversion of the polymer to an amorphous ceramic material, as shown in [10,12,13]:

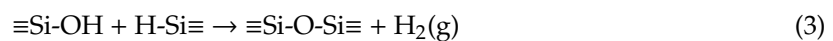


The mass loss is decreasing to  $-3.5$  wt.% in the (24 Si + 12 B) vol.% filled system due to reduction of the PHPS content. The chemical rearrangement of the constituents of the  $\text{Si}_x\text{N}_y$  ceramic in the nascent state takes place above  $500$  °C.



**Figure 4.** Mass change curves of crosslinked plain PHPS and PHPS + (24Si + 12 B) volume% of fillers (balance to 100 vol.% PHPS) after pyrolysis in  $\text{N}_2$  atmosphere.

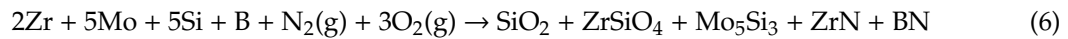
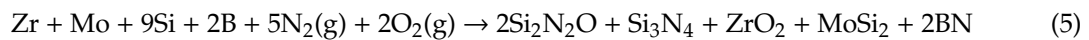
Due to the presence of hydroxyl groups on the filler's surface [14], and the high reactivity and sensitivity of PHPS to moisture and  $-\text{OH}$  groups, siloxane bonds ( $\text{Si}-\text{O}-\text{Si}$ ) form according to chemical reactions (3) and (4), [10]:



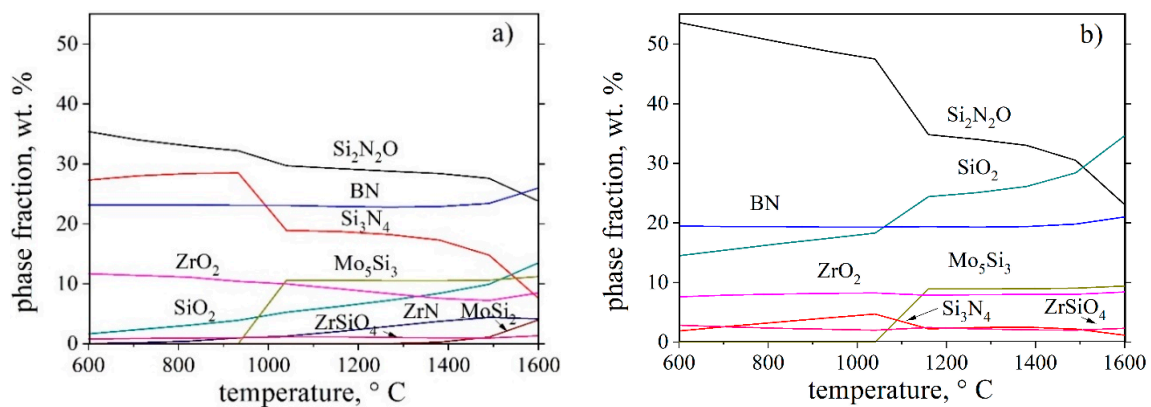
At this stage, the nitrogen atoms are particularly replaced by oxygen atoms, thus, the formation of silicon oxide and silicon oxynitride amorphous ceramic besides of  $\text{Si}_3\text{N}_4$  might be possible [22]. Günthner et al. [11,15] reported nanoscopic free Si resulting from the process of PHPS polymer-to-ceramic conversion. The interaction of this fine, particular Si with  $\text{N}_2$  atmosphere might result in an increase of the  $\text{Si}_3\text{N}_4$  or  $\text{SiNO}$  phase amount, which explains the mass gain observed from  $750$  °C to  $1000$  °C as detected at the PHPS/filler curve (see Figure 4). The total ceramic yield of the composite after thermal treatment at  $1000$  °C in  $\text{N}_2$  is 96.6 wt.%.

To design a possible coating composition as the result of PHPS interaction with filler materials, pyrolysis atmosphere and alloy as substrate, thermodynamic calculations of the most probable chemical reactions were carried out. The ability of the reactants to interact and the directions of the chemical reactions were determined by a change of the isobaric-isothermal potential,  $\Delta G_{\circ T}$  [24]. For the equilibrium calculation, the amounts of Si,  $\text{N}_2$ ,  $\text{H}_2$  as components of PHPS were set to 100 g of precursor (see Table 1), and Si and B as fillers were added according to their mass fraction in the crosslinked samples. The amount of Mo, Zr and B from the alloy surface was determined by "trial and error", choosing the minimal quantity of component for the formation of expectable phases without their excess as pure elements at the phase equilibrium. The amount of  $\text{O}_2$  was varied in the range from 0.001 to 0.1 mol. The amount of  $\text{N}_2$  was varied in order to investigate its influence on the phase formation within the reaction system.

With respect to thermodynamics, the formation of  $\text{Si}_3\text{N}_4$ ,  $\text{Si}_2\text{N}_2\text{O}$  and  $\text{SiO}_2$  may occur according to chemical reactions shown in Equations (5) and (6), ( $\Delta G_{1000\text{ }^\circ\text{C}} = -2648.294\text{ kJ}$  and  $\Delta G_{1000\text{ }^\circ\text{C}} = -2942.361\text{ kJ}$ , respectively):



Besides, at the equilibrium, other phases based on the above listed elements and their compounds are present: BN,  $\text{ZrO}_2$  and some amount of  $\text{Mo}_5\text{Si}_3$ ,  $\text{MoSi}_2$ ,  $\text{ZrSiO}_4$  and  $\text{ZrN}$ . The presented phases were found simultaneously at the equilibrium in the investigated temperature range between  $600\text{ }^\circ\text{C}$  and  $1600\text{ }^\circ\text{C}$ , where  $\text{Mo}_5\text{Si}_3$  and  $\text{MoSi}_2$  were found to be likely at higher temperatures, here starting from  $956\text{ }^\circ\text{C}$  and  $1220\text{ }^\circ\text{C}$ , respectively, see Figure 5. Molybdenum silicides and zirconium phases in the ceramic-metal system may be chemically compatible with the Mo-14.8Zr-26B alloy and being present in a small quantity may lead to a minimization of thermal mismatch between the substrate and the coating, thereby proving a good adhesion. On the other hand,  $\text{ZrO}_2$  and  $\text{ZrN}$  phases are known for their promising chemical and mechanical properties, which promulgate them as perspective candidates with high wear and corrosion resistance [25].



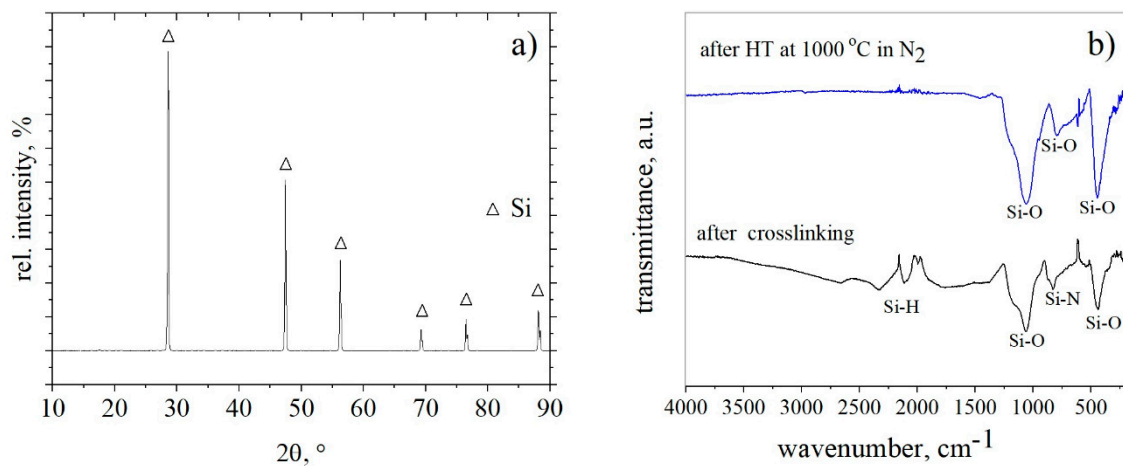
**Figure 5.** Calculated equilibrium phase amounts in the temperature range from  $600\text{ }^\circ\text{C}$  to  $1600\text{ }^\circ\text{C}$  of (a) Si-N<sub>2</sub>-H<sub>2</sub>-O<sub>2</sub>-Mo-Zr-B (0.238-0.596-0.193-0.050-0.082-0.019-0.036) mol and (b) Si-N<sub>2</sub>-H<sub>2</sub>-O<sub>2</sub>-Mo-Zr-B (0.238-0.596-0.193-0.100-0.082-0.019-0.036) mol.

The main compound in the coating system is  $\text{Si}_2\text{N}_2\text{O}$ . In accordance with the above calculated thermodynamic systems (see Figure 5a), the amount of silicon oxynitride is 24 wt.% to 36 wt.% at temperatures between  $600\text{ }^\circ\text{C}$  and  $1600\text{ }^\circ\text{C}$ . The increase of  $\text{O}_2$  in the system from 0.05 mol to 0.1 mol affects the formation of both,  $\text{Si}_2\text{N}_2\text{O}$  and  $\text{SiO}_2$ , (Figure 5b). In a temperature range between  $1000\text{ }^\circ\text{C}$  and  $1100\text{ }^\circ\text{C}$ , a noticeable decrease of the  $\text{Si}_2\text{N}_2\text{O}$  phase amount and an increase of  $\text{SiO}_2$  phase quantity were found. The maximal amount of  $\text{Si}_3\text{N}_4$  in this system was calculated for a temperature of  $1050\text{ }^\circ\text{C}$ . According to the previous diagrams on the assumption of  $\text{O}_2$  amount, the correlation of a maximum quantity of  $\text{Si}_2\text{N}_2\text{O}$  and  $\text{SiO}_2$  protective phases and a minimum quantity of  $\text{Mo}_5\text{Si}_3$  and  $\text{MoSi}_2$  accessory phases, which are undesired in a high amount in the coating, is evident in the temperature range of  $900\text{--}1100\text{ }^\circ\text{C}$ . Thus, this temperature interval should be chosen for the coating's pyrolysis.

It was found in our calculations that an increase of the  $\text{N}_2$  amount within the system does not influence the equilibrium of the solid phases. From this, it was concluded that the amount of  $\text{N}_2$  resulting from the PHPS preceramic polymer is sufficient to form nitrogen-containing phases within these systems, and pyrolysis in argon instead of in nitrogen becomes an option.

### 3.2.2. Phase Composition

To estimate the real phase composition of the coatings after processing, XRD and FT-IR analysis were performed with the coating materials. The coating slurry was dried and crosslinked in a drying furnace at 110 °C for 1.5 h followed by grinding with a mortar and pestle to a fine powder (<44 μm) and heat-treated in a tubular furnace in N<sub>2</sub> at 1000 °C. The powders obtained after pyrolysis were subjected to X-ray diffraction analysis, (Figure 6a).



**Figure 6.** (a) X-ray diffraction patterns of the coating material after pyrolysis at 1000 °C in N<sub>2</sub> (heating rate and cooling rate were 3 K/min and the dwell time at max temperature was 1 h); (b) FT-IR (Fourier-transform infrared spectroscopy) spectra of the plain, Si- and B-filled coating system after crosslinking at 110 °C in air and after pyrolysis at 1000 °C in nitrogen.

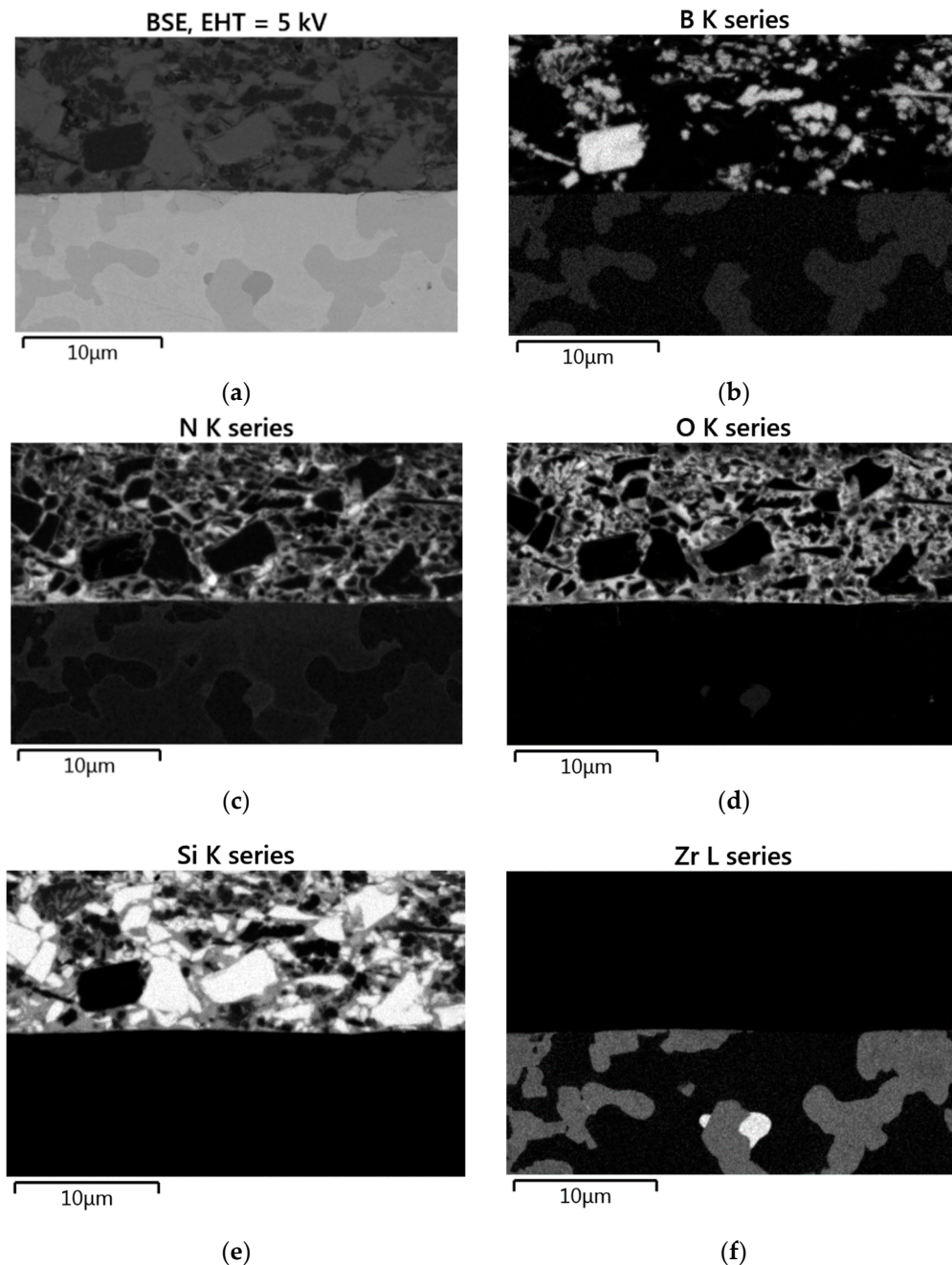
The only crystalline phase in the coating system was silicon; silicon and boron were added as fillers, (see coating composition in Materials and methods). Boron was not detected by XRD analysis because of its amorphous state.

In the FT-IR spectra, the crosslinked absorption bands were detected corresponding to Si-H stretching vibration at 2120 cm<sup>-1</sup>, Si-N vibrations at 840–1020 cm<sup>-1</sup> and Si-O vibrations at 1063 cm<sup>-1</sup> (stretching mode region) and at 456 cm<sup>-1</sup> (rocking mode region) [26–28], (Figure 6b). This demonstrates that at the stage of crosslinking, oxygen atoms particularly replaced nitrogen atoms resulting in silicon oxynitride and silicon oxide amorphous ceramic. After pyrolysis in nitrogen at 1000 °C, the Si-H absorption band completely disappears because of the dehydrogenation reaction, and the only absorption peaks of Si-O bonds were detected in the composite. The Si-O vibrational peak at 1063 cm<sup>-1</sup> was found to have a higher intensity compared to that of the crosslinked material. From this, we conclude a higher number of Si-O-Si units within the pyrolyzed material. The incorporation of nitrogen into the material cannot be excluded nor confirmed. Since no other phase than elemental silicon was found by XRD, the SiO<sub>2</sub> phase is supposed to be an amorphous/glassy phase.

This coincides well with the results of Günthner et al. [11,15] as mentioned above. Free Si might act as an oxygen trap and provide a self-healing functionality due to the formation of SiO<sub>2</sub>. In summary, the coating system consists of an amorphous SiO<sub>2</sub> phase, a SiON phase which is probably Si<sub>2</sub>ON<sub>2</sub>, elemental silicon and elemental boron.

In another set of experiments, real coatings on Mo-14.8Zr-23B and on Mo-14.8Zr-26B were characterized after pyrolysis in nitrogen at 1000 °C. Scanning electron microscope (SEM) analyses and EDS element mapping results are shown exemplarily for the Mo-14.8Zr-23B sample in Figure 7.





**Figure 7.** (a) SEM micrograph, and (b–f) EDS element mapping of a coated Mo-14.8Zr-23B alloy after pyrolysis at 1000 °C in N<sub>2</sub>.

The coating layers did not show any phase segregation across the entire samples cross section, and its average thickness on the planar sample surfaces was 400 μm and 420 μm for Mo-14Hf-23B and Mo-14.8Zr-26B, respectively. However, it should be noted that the thickness of the coating was only 10 μm at the corners of the samples, while in some other regions (typically at the horizontal flat surfaces) the coating was more than 600 μm thick.

Element mappings confirmed the presence of Si, N, B and O in the coating layer. The light white particles of Si are surrounded by the PDC-matrix appearing light grey, and particularly reacted to form SiO<sub>x</sub> or SiO(N) phases. The close neighborhood of boron und silicon suggests the formation of a borosilicate glassy phase. This glassy phase might have a viscosity low enough to fill gaps and

cracks in the substrate material and protect it from oxidation: oxygen possesses a very small diffusion coefficient ( $\approx 1 \times 10^{-9} \text{ cm}^2/\text{s}$ ) when passing  $\text{SiO}_2\text{-B}_2\text{O}_3$  glass [29].

### 3.2.3. Cyclic Oxidation of the Coated Samples

The coated Mo-14Hf-23B and Mo-14.8Zr-26B alloys were exposed to cyclic oxidation at  $800^\circ\text{C}$  in air; the corresponding mass change curves are shown in Figure 1 in comparison to the non-coated samples. During the first 10 h of the oxidation process, a mass increase of the coated samples was detected. This behaviour is opposite to the one observed for the uncoated specimens, which show a significant mass loss already at the beginning of the oxidation test. The mass gain of the coated Mo-14Hf-23B and Mo-14.8Zr-26B alloys amounts  $13.2 \text{ mg}/\text{cm}^2$  and  $21.5 \text{ mg}/\text{cm}^2$ , respectively, after the first 10 h of exposure. This might be explained by the formation of a glassy oxide layer on the sample surface resulting from silicon and boron oxidation, as indicated by EDS element mappings as shown in Figure 8. After 10 h of oxidation the growth of the protective oxide layer reaches a steady state. This is indicated by stagnation in the mass gain.

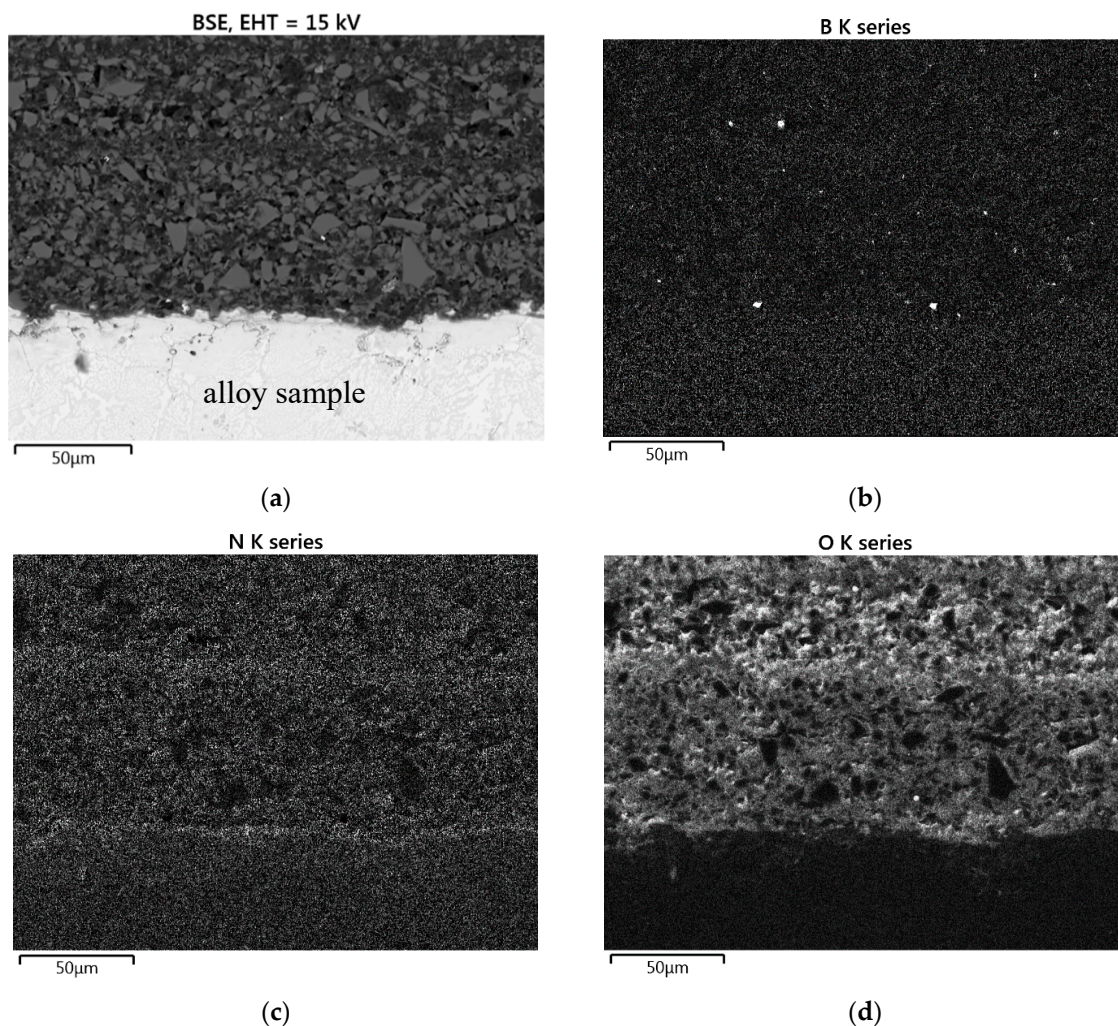
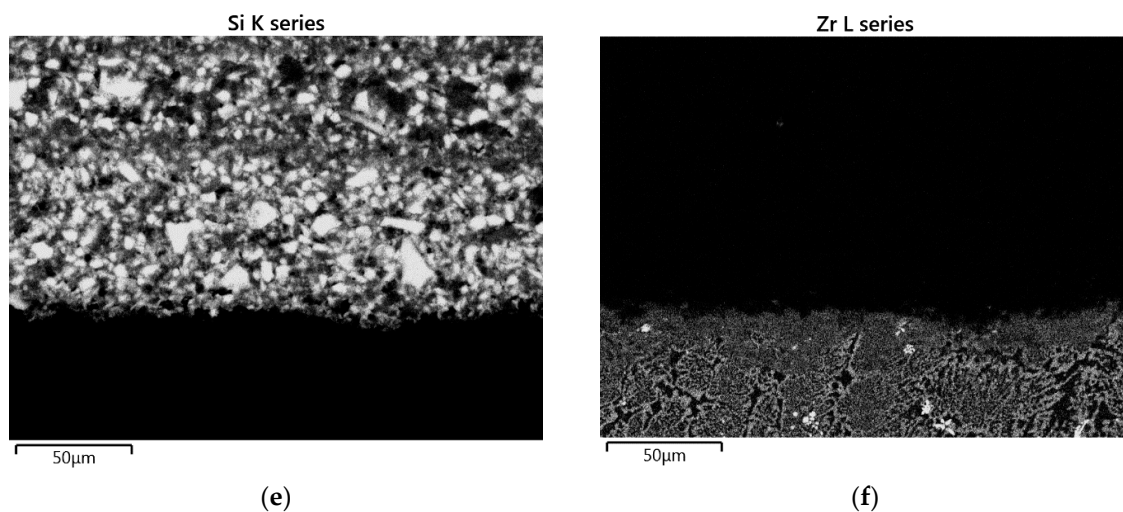


Figure 8. Cont.



**Figure 8.** (a) SEM micrograph, and (b–f) EDS elements mapping of a PHPS/Si-B-coated Mo-14.8Zr-26B alloy after crosslinking, with pyrolysis at 1000 °C in nitrogen and oxidation testing at 800 °C in air.

Even after 50 h of oxidation testing in air at 800 °C, the distribution of Si, N, B and O in the oxidized coating layers are retained (EDS mapping results, Figure 8), and no hints to the presence of Zr or Mo oxides was found in the coating layers. This points out that no oxidation processes occur after the alloy samples were coated, crosslinked, pyrolyzed and exposed to air. However, there was a slight increase of the coating thickness after high temperature exposure to air; coating thicknesses increased to 460 µm, and 480 µm for Mo-14Hf-23B and Mo-14.8Zr-26B alloy, respectively. This is caused by a SiO<sub>2</sub> surface layer formation resulting from the chemical reaction of free Si with oxygen, see Figure 8.

Thus, the coating system with PHPS + (24Si + 12B) vol.% provides an excellent protection effect against oxidation at 800 °C.

#### 4. Conclusions

Cyclic oxidation tests of Mo-14Hf-23B and Mo-14.8Zr-26B alloys demonstrated a poor oxidation resistance of this type of materials at 800 °C in air. This is caused mainly by catastrophic Mo oxidation at intermediate temperatures. However, a new coating system was developed based on the preceramic polymer PHPS with Si and B fillers, which was applied to the substrate materials to protect the alloys from oxidation. After crosslinking and pyrolysis, the PHPS polymer transformed mainly into SiO<sub>2</sub> and SiO (N) amorphous glasses/ceramics and dispersed elemental Si; Si is a precursor to form a dense and stable SiO<sub>2</sub> coating on Mo-(Hf,Zr)-B substrates. The PHPS/Si-B coating provides excellent protection against oxidation at 800 °C in air, which is the harshest temperature for uncoated alloys of this type.

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