Diferrocenyl Thioketone: Reactions with (Bisphosphane)Pt(0) Complexes—Electrochemical and Computational Studies

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Abstract: Diferrocenyl thioketone reacts smoothly with (bisphosphane)Pt(0) complexes in toluene solution at room temperature yielding 1:1 adducts identified as ferrocenyl (Fc) functionalized platinathiiranes. Their structures were unambiguously confirmed by means of spectroscopic methods as well as by X-ray diffraction analysis. A unique, ferrocene-rich platinathiirane, bearing three Fc-units, was prepared starting with [bis(diphenylphosphino)ferrocene] Pt(0)(η²-norbornene). For comparison, a similar platinathiirane with one Fc-unit was obtained from the reaction of the latter complex with thiobenzophenone. Quantum-chemical calculations were carried out to describe the bonding pattern and frontier molecular orbitals of the ferrocene-rich platinathiirane complexes. These calculations confirmed that the C=S bond loses its formally double-bond character upon complexation (bisphosphane)Pt(0). Cyclic voltammetry measurements were performed to characterize the obtained platinathiiranes in CH₂Cl₂ solutions. For comparison, the cyclic voltammogram for diferrocenyl thioketoneas a mixed-valent (Fe²⁺-Fe³⁺) compound was also recorded and analyzed. The results point out to a diffusion controlled electrode process in case of diferrocenyl thioketone and mixed diffusion and adsorption controlled electrode process in the case of the studied platinathiiranes.

Keywords: thioketones; ferrocene; Pt(0) complexes; platinathiiranes; cyclic voltammetry; DFT calculations

1. Introduction

In our continuing studies on the chemistry of thiocarbonyl compounds, we demonstrated that in contrast to information available in some earlier monographs [1], reviews [2], and textbooks [3], aryl, hetaryl, and ferrocenyl thioketones should currently be considered as versatile starting materials for the synthesis of more complex sulfur-containing as well as sulfur-free organic compounds [4]. In this context, their 1,3-dipolar cycloadditions and hetero-Diels-Alder reactions leading to five- or six-membered sulfur heterocycles, correspondingly, are of special importance. Because of the high reactivity in these reactions, thioketones were named as superdipolarophiles [5] and superdienophiles [6], respectively. It is well established that stability of thioketones and consequently their synthetic utility depend
mixed anhydride” is a method of choice to get to dispiro-cycloaliphatic substituted 1,2,4-trithiolanes, the initiating step is the oxidative insertion of thioketones [16]. These studies showed that the M(CO)₅ to 2,2-diferrocenyl 3,3-diphenylthiirane [9]. In earlier reports, M(CO)₅ and the reaction with diphenyl diazomethane led after two-fold elimination of nitrogen [N]. The reaction with 2,3-dimethyl-buta-1,3-diene yielded the corresponding 3,6-dihydro-2-thiapyran [13] and the reaction with diphenyl diazomethane led after two-fold elimination of nitrogen [N₂] and sulfur [S] to 2,2-diferrocenyl 3,3-diphenylthiirane [9]. In earlier reports, M(CO)₅ complexes (M = Cr, W) of 1a were obtained and their structures were studied by means of spectroscopic and X-ray diffraction (XRD) methods [14,15]. These studies showed that the M(CO)₅ moiety is coordinated via the sulfur atom of the C=S group.

In a series of our recent publications we described the formation of platinathiiranes 3 in reactions of [Pt(0)L₂(η²-olefine)] (L = phosphanes) 2 (Figure 1) with diverse aromatic and cycloaliphatic thioketones [16].

![Scheme 1](image)

**Scheme 1.** Preparation of thioketone 1a by thionation of diferrocenyl ketone using Lawesson’s reagent (LR) [9]. TFAA/TFA/CH₂Cl₂ denotes trifluoroacetic anhydride in the presence of trifluoroacetic acid in solution.

In spite of the fact, that ferrocene itself is named as an exceptional molecule [12] and organic compounds functionalized with the electroactive ferrocenyl moiety are of great importance in diverse field of applied chemistry (materials chemistry, medicinal chemistry, electrochemistry), thioketone 1a has been little explored in organic and coordination chemistry as a useful substrate for synthesis of such compounds and only few reports are known regarding data on its reactivity, e.g., thia-Diels-Alder reaction with 2,3-dimethyl-buta-1,3-diene yielded the corresponding 3,6-dihydro-2H-thiapyran [13] and the reaction with diphenyl diazomethane led after two-fold elimination of nitrogen [N₂] and sulfur [S] to 2,2-diferrocenyl 3,3-diphenylthiirane [9]. In earlier reports, M(CO)₅ complexes (M = Cr, W) of 1a were obtained and their structures were studied by means of spectroscopic and X-ray diffraction (XRD) methods [14,15]. These studies showed that the M(CO)₅ moiety is coordinated via the sulfur atom of the C=S group.

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![Figure 1](image)

**Figure 1.** Bidentate phosphanes used as ligands in platinum complexes 2a–c. L₂ = 1,2-bis(diphenylphosphino)benzene (dppe), 1,8-bis(diphenylphosphino)naphthalene (dppn), 1,1'-bis(diphenylphosphino)ferrocene (dpff); nbe = norbornene.

The same products were also obtained by treatment of 1,2,4-trithiolanes with 2. In the case of dispiro-cycloaliphatic substituted 1,2,4-trithiolanes, the initiating step is the oxidative insertion of
the L\textsubscript{2}Pt(0) complex fragment into the sulfur–sulfur bond of the 1,2,4-trithiolane ring. Subsequent (3 + 2)-cycloelimination of the intermediate formed thereby, releases the corresponding thioketone which in the next step interacts with another equivalent of 2 forming platinathiiranes 3. In these multi-step reactions, the second type of products formed, was identified as four membered (dithiolato)platinum(II) complexes 4 [17].

On the other hand, in reactions of 3,3,5,5-tetraphenyl 1,2,4-trithiolane (5) another mechanism comprising the in situ generated diphenyldithiirane/thiobenzophenonen-S-sulfide (6/7) and thiobenzophenone (1b), which are both formed upon thermal (3 + 2)-cycloreversion of 5, was postulated. These components are trapped with a phosphine complex of type 2 yielding side by side the corresponding platinathiiranes 3 and dithiolato complexes 4, respectively. Complex 3 was also formed by direct reaction of 2 with 1b (Scheme 2) [18].

![Scheme 2](image_url)

**Scheme 2.** Preparation of platinathiiranes 3 and dithiolato complexes 4 in reactions with 1,2,4-trithiolane 5.

Reactions of 1,2,4-trithiolanes and their S-oxides with Pt(0) complexes leading to diverse platinathiiranes and dithiolato complexes of type 3 and 4, respectively, are summarized in a recent review [19].

The goal of the present work is the synthesis and characterization of novel platinathiiranes of type 3 derived from thioketone 1a and diverse (bisphosphane)Pt(0) complexes of type 2 presented in Figure 1. Furthermore, physico-chemical properties and structures of the ferrocenyI functionalized platinathiiranes 3 should be studied by means of electrochemical and computational methods.

## 2. Materials and Methods

### 2.1. Preparation and Identification of Platinathiiranes 3a–d

**Materials and Techniques.** The \textsuperscript{1}H and \textsuperscript{13}C[\textsuperscript{1}H] spectra were recorded with a Bruker Avance 200 MHz spectrometer (Terre Haute, IN, USA). Chemical shifts are given in parts per million with reference to internal SiMe\textsubscript{4} or CHCl\textsubscript{3}. The mass spectrum was recorded with a Finnigan MAT SSQ 710 instrument (Auburn, CA, USA). The IR spectra were measured with a Perkin–Elmer System 2000 FT-IR spectrometer (Waltham, MA, USA). Elemental analysis was performed with a Leco CHNS-932 apparatus (Madrid, Spain). Silica gel 60 (0.015–0.040 mm) was used for column chromatography, and thin layer chromatography (TLC) was performed by using Merck (Kenilworth, NJ, USA) TLC aluminum sheets (Silica gel 60 F254). Solvents from Fisher Scientific (Ottawa, Ontario, Canada) and other chemicals from Acros were used without further purification. All solvents were dried and distilled prior to the use according to standard methods. The starting materials were prepared according to literature procedures [9,18,20].

**General procedure:** To a stirred solution of 0.063 mmole (bisphosphane)Pt(nbe) 2a–c [20] in 15 mL toluene one equivalent of thioketone 1a (or 1b) in toluene (10 mL) was added dropwise, whereupon
the color of the solution changed. The mixture was stirred overnight at room temperature. Then, the solvent was removed under reduced pressure at room temperature and the solid residue was washed three to four times with diethyl ether. Next, the residue was dissolved in a small amount of toluene and subsequent diffusion of n-pentane into this concentrated solution gave crystals of complexes 3a–d suitable for SC-X-ray diffraction. Structures of the synthesized complexes 3 are shown in Figure 2.

Figure 2. Platinathiranes 3a–d investigated in this work.

(dppbe)Pt(2-diferrocenylthioketone) 3a. Orange colored crystals, yield: 61 mg (91%).1H (CD2Cl2, 400 MHz): 7.75 (m, 1H, dppbe), 7.65 (m, 5H, dppbe), 7.47 (m, 2H, dppbe), 7.39 (m, 6H, dppbe), 7.33 (m, 2H, dppbe), 7.23 (m, 8H, dppbe), 3.92 (m, 12H, 10H C5H5 and 2H C5H4), 3.85 (m, 2H, C5H4), 3.82 (m, 2H, C5H4), 3.75 (m, 2H, C5H4).13C{1H} (CD2Cl2, 101 MHz): 135.7 (s), 134.8–134.4 (m), 130.4 (m), 129.8 (m), 129.3 (m), 128.4 (m), 127.9 (m), 125.5 (m), 128.4 (m), 127.7 (m), 100.3 (m), 76.1 (m), 74.7 (m), 72.9 (m), 71.4 (m), 70.9 (m), 70.3 (m), 69.9 (m), 69.4 (m), 66.1 (m), 65.8 (m). 31P{1H} (CD2Cl2, 162 MHz): 4.63 (d with 195Pt sat., 1JPPt = 2835.0, 2JPPt = 28.1 Hz), 36.52 (d with 195Pt sat., 1JPPt = 4167.2, 2JPPt = 28.0 Hz).195Pt{1H} (CD2Cl2, 129 MHz): –5059 (dd, 1JPPt = 4168 Hz, 2JPPt = 2835 MHz). ESI-MS(+mode) m/z: 1054.8 [M-2H]+ (including isotope pattern as calculated). EA: C51H42Fe2P2PtS-0.5 toluene-0.25 pentane (calc./found): C: 59.80/59.71; H: 4.41/4.40; S: 2.86/2.65.

(dppn)Pt(2-diferrocenylthioketone) 3b. Orange colored crystals, yield: 47 mg (68%).1H (CD2Cl2, 400 MHz): 8.05 (m, 2H, o-H of dppn bridge), 7.42 (m, 2H, CHar), 7.33 (m, 2H, CHar), 7.25 (m, 8H, o-∥-C6H5), 7.14 (m, 8H, o-∥-C6H5), 6.91 (m, 4H, CHar), 6.58 (CHar), 4.20 (m, 10H, Cp C5H5), 4.12 (m, 2H, Cp C5H4), 3.98 (m, 2H, Cp C5H4), 3.90 (m, 2H, Cp C5H4), 3.85 (m, 2H, Cp C5H4).13C{1H} (CD2Cl2, 101 MHz): 134.0 (m), 133.6 (m), 133.3 (m), 130.0 (m), 129.6 (m), 129.6 (m), 128.3 (m), 127.9 (m), 125.5 (m), 70.3 (m), 69.4 (m), 66.1 (m), 65.8 (m). 31P{1H} (CD2Cl2, 162 MHz): 4.63 (d with 195Pt sat. 1JPPt = 2704 Hz, 2JPPt = 22.5 Hz), 4.26 (d with 195Pt sat. 1JPPt = 3970 Hz, 2JPPt = 22.5 Hz). ESI-MS(+mode) m/z: 1104.6 [M-H]+ (including isotope pattern as calculated minus H). EA: C55H44Fe2P2PtS-0.5 toluene-0.25 pentane (calc./found): C: 59.74/59.15; H: 4.01/3.95; S: 2.90/2.62.

(dppf)Pt(2-diferrocenylthioketone) 3c. Orange colored crystals, yield: 41 mg (56%).1H (CD2Cl2, 400 MHz): 8.01 (m, 4H, C6H5), 7.5 (m, 6H, C6H5), 7.24 (m, 2H, C6H5), 7.03 (m, 8H, C6H5), 4.34 (m, 4H, C6H5), 4.19 (m, 12H, C6H5), 4.09 (m, 2H, C6H5), 3.83 (m, 4H, C6H5), 3.74 (m, 2H, C6H5), 3.66 (m, 2H, C6H5).13C{1H} (CD2Cl2, 101 MHz): 135.7 (s), 134.8–134.4 (m), 130.4 (m), 129.8 (m), 129.3 (m), 128.4 (m), 127.7 (m), 100.3 (m), 76.1 (m), 74.7 (m), 72.9 (m), 71.4 (m), 70.9 (m), 70.3 (m), 69.9 (m), 69.5 (m), 69.3 (m), 65.8 (m). 31P{1H} (CD2Cl2, 162 MHz): 17.66 (s with 195Pt sat. 1JPPt = 4599 Hz), 14.14 (s with 195Pt sat. 1JPPt = 3970 Hz). 195Pt{1H} (CD2Cl2, 129 MHz): –4900 (dd, 1JPPt = 4597 Hz,
1JPP = 3117 MHz). ESI-MS((+mode) m/z: 1161.6 [M-2H]+ (including isotope pattern as calculated minus 2H). EA: C35H44Fe3P2PtS0.25 toluene (calc./found): C: 57.44/57.76; H: 4.08/4.11; S: 2.70/2.76.

(dppf)Pt(2-thiobenzophenone) 3d. Colorless crystals, yield: 51 mg (85%). 1H (600 MHz, CD2Cl2): 7.91–7.83 (m, 4 H, CHar), 7.45 (m, 6 H, CHar.), 7.29–7.20 (m, 6 H, CHar), 7.05 (m, 8 H, CHar), 6.92 (m, 4 H, CHar), 6.89–6.85 (m, 2 H, CHar), 4.16 (m, 12 H, C5H4 of cp). 13C1H) (151 MHz, CD2Cl2): 134.7 (C9), 134.7 (C9), 134.4 (C9), 134.4 (C9), 130.5 (C9), 130.1 (C9), 128.5 (C9), 128.4 (C9), 127.9 (C9), 127.1 (C9), 124.2 (C9), 75.9 (m, C5P, 75.0 (s, C5P), 73.0 (m, C5P), 72.1 (s, C5P). 31P1H) (243 MHz, CD2Cl2): 18.82 (s with 195Pt-sat 1JPP = 4508 Hz), 17.64 (s with 195Pt-sat 1JPP = 3135 Hz) P-P coupling not resolute. 195Pt1H) (129 MHz, CD2Cl2): -4988 (dd, 1JPP = 4511, 3132 Hz). MS (DEI): m/z (%) = 947 [M+] (8), 749 [M+-thioketone] (28), 198 [thioketone+] (98), 186 [ferrocene+] (99), 185 [diphenylphosphane+] (99), 165 [thioketone+-S] (100), 121 [thioketone+-phenyl] (100), 108 [phenylphosphane+] (99), 77 [phenyl+] (99). EA: C47H38Fe2P2PtS0.5CD2Cl2 (calc./found): C: 57.56/57.80; H: 4.07/4.30; S: 3.23/3.60.

2.2. Crystal Structure Determination

The intensity data for the compounds were collected on a NoniusKappaCCD (Bruker, Austin, TX, USA diffractometer using graphite-monochromated Mo-Kα radiation. Data were corrected for Lorentz and polarization effects; absorption was taken into account on a semi-empirical basis using multiple-scans [21–23]. The structures were solved by direct methods (SHELXS [23]) and refined by full-matrix least squares techniques against F2 by the SQUEEZE routine of the program PLATON [26] resulting in 40 electrons

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2.3. Electrochemical Measurements

Cyclic voltammetry (CV) measurements were conducted using the three-electrode technique, with glassy carbon disk (diameter d = 1.6 mm) as the working electrode, Ag/Ag+ in acetonitrile (MeCN) as the reference electrode, and Pt wire as the counter electrode, with a Reference 600 Potentiostat (Gamry Instruments). Corrections for the IR drop were performed for all experiments. All experiments were performed in CH2Cl2 solutions (concentration of the complexes 1.0 mmol/L) containing 0.1 mol/L [nBu4N][BF4] at room temperature. Solutions were deaerated by N2 purge for 5–10 min, and a blanket of N2 was maintained over the solutions during the measurements. The glassy carbon disk was polished on a felt tissue with alumina before each measurement. All potential values reported in this paper are referenced to the potential of the ferrocenium/ferrocene (Fc+/Fc) couple. The dependence of the peak currents (I) from the square root of scan rate (v1/2) and logarithm of peak currents (log I) versus logarithm of scan rates (log v) for complexes 3a–d (1.0 mmol/L) in CH2Cl2 - [nBu4N][BF4] (0.1 m/L) are given in Figure S1.
2.4. Theoretical Calculations

The structure of complex 3c was optimized at the BP-D/def2-TZVP level of theory [28–35] (Tables S3 and S4). The XRD geometry of 3c was used as a starting point for the optimization (Figure S2). Vibrational frequency calculations confirmed that the optimized geometry corresponded to a true minimum on the potential energy surface of 3c. The same calculations were also carried out for the separate 1a and (dppf)Pt fragments constituting 3c. All these calculations were done with the TURBOMOLE 7.2 program [36]. For the optimized complex, a single-point energy calculation with a slightly larger basis set (def2-TZVPP) [34] was performed using Gaussian 09 D.01 [37]. This program allowed to generate the molecular wave function files required to analyze the bonding situation in terms of natural bond orbitals (NBO) [38] and the quantum theory of atoms in molecules (QTAIM) [39]. The NBO analysis was done with the NBO 3.1 subroutine [40] incorporated into Gaussian 09 D.01. This subroutine was also used by Multiwfn 3.5 [41] to express the composition of frontier molecular orbitals in terms of natural atomic orbitals [42]. In the case of QTAIM, its implementation available in AIMAll 14.11.23 [43] was utilized. To verify the results obtained from BP-D/def2-TZVPP, an additional single-point energy calculation was carried out using the popular B3LYP hybrid density functional [44–46] combined with the def2-TZVPP basis set (these results are presented in Tables S3 and S4 and in Figures S3 and S4). Results from both density functionals turned out to be in qualitative agreement, and therefore, only the BP-D results will be quoted in Section 3.3. Cartesian coordinates (in Å) for 3c, 1a, and (dppf)Pt-fragment are collected in Tables S5–S7, respectively (the SI part, see S4).

3. Results and Discussion

3.1. Synthesis and Characterization of Platinathiiranes 3a–d

Acetylation of ferrocene with ferrocenyl carboxylic acid using the so-called “mixed anhydride” prepared in situ by treatment of the latter with trifluoroacetic anhydride (TFAA) opens a straightforward access to the required diferrocenyl ketone. Subsequent thionation of this ketone with Lawesson’s reagent in toluene solutions leads to the desired thioketone 1a in high yield [9].

The test experiment for the preparation of platinathiiranes was performed starting with 1a and 2a in toluene solution at room temperature using equimolar amounts of both compounds. After 24 h the solvent was removed and a crystalline product was obtained. The $^{31}$P{H} NMR spectroscopic analysis showed a set of signal characteristic for aromatic CH units as well as for two chemically equivalent ferrocenyl groups.

In the $^{13}$C{H} NMR spectrum the signal located at 102.9 ppm confirmed the presence of the C–S moiety. In comparison with the resonance signal of 1a (237.1 ppm) the $^{13}$C signal of the thiocarbonyl group, was drastically shifted to lower frequencies. Finally, the registered $^3\text{J}({\text{PtP}})$ = 2835 and 4167 Hz, respectively. In addition, the ESI-MS showed the m/z = 1054.8 which corresponds to the calculated [M–2H]$^+$ ion for the expected 1:1 complex 3a (Scheme 3).

![Scheme 3. Synthesis of the platinathiirane complex 3a derived from 1a and 2a.](image)
The obtained product 3a was crystallized by diffusion of pentane into the concentrated toluene solution and single crystals suitable for XRD analysis were obtained. The analysis confirmed unambiguously the proposed structure of platinathiirane 3a (Figure 3).

![Molecular structure (50% probability) of complex 3a. Hydrogen atoms, solvent molecule (toluene), as well as the phenyl rings of the dppbe ligand are omitted for clarity.](image)

The bond lengths Pt-S(1) [2.3014(8) Å], S(1)-C(1) [1.791 (3) Å], and Pt-C(1) [2.155(3) Å] (see Table S1) are comparable with values reported for analogous platinathiiranes containing the same bisphosphane ligand [18,47]. The same method was used for the preparation of platinathiiranes 3b and 3c starting with 1a and Pt(0) complexes 2b and 2c, respectively. The structures of complexes 3b and 3c were also established by means of spectroscopic methods as well as by XRD analysis (Figure 4).

![Molecular structures (50% probability) of complexes 3b (left) and 3c (right). Hydrogen atoms, solvent molecule (CH2Cl2) (in 3b) as well as the phenyl rings of the dppn (3b) and dpff (3c) ligands are omitted for clarity.](image)

In extension of the study, reaction of 1b with complex 2c was also carried out and the expected platinathiirane 3d was isolated in 85% yield and collected spectroscopic data confirmed unambiguously its structure (see Experimental).

### 3.2. Electrochemical Studies

In the second part of the study, the obtained platinathiiranes 3a–d were investigated by CV in CH2Cl2 solutions. For comparison, the cyclic voltammogram of thioketone 1a was also recorded and analyzed (Figures 5 and 6). Cyclic voltammograms of complexes 3a–d are shown in Figures 7–10.
3.2.1. Thioketone 1a

The electrochemical behavior of biferrocene can be affected by at least two types of interaction between the two ferrocene moieties \[48\]. The first is the interaction through the bridge atoms, and the second is the interaction between two iron–iron centers.

Biferrocene undergoes two reversible, one-electron oxidation processes with the separation potential value of 340 mV \[49,50\]. These two distinct one-electron oxidations can be explained by the fact that cyclopentadienyl rings around the Fe(1) and Fe(2) are differently organized. The rings around the Fe(1) atom are eclipsed, whereas these around Fe(2) are staggered. Because of this orientation the Fe–cyclopentadienyl rings mean bond lengths are different in case of Fe(1) and Fe(2) which results in electronic delocalization and the biferrocenium ion can be classified as a mixed-valent (Fe\(^{II}\)-Fe\(^{III}\)) compound displaying localized charge. Both Fe atoms are in different oxidation states, Fe(II) and Fe(III), respectively. Thus the first oxidation process removes an electron from the Fe(II).

In diferrocenylmethane the separation potential value is of 120 mV \[51\]. This may point to less communication between the two ferrocene moieties. But if we look at the two-bridged [1,1]-ferrocenephane, the separation between oxidation potentials is about twice as large, 190 mV \[48\]. The authors assume that difference in that case is due to a direct Fe–Fe electrostatic field interaction which may increase the communication in case of [1,1]-ferrocenephane where the rotation around –CH\(_2–\) function is suppressed. This assumption can be also supported by the structure in which both of the iron atoms are placed off-center in the ferrocene moieties. In such configuration the intramolecular repulsion of the nonbonding electrons of iron atoms can be reduced \[52\].

1,2-Diferrocenylethane displays in dichloromethane solution a single two-electron oxidation process which is reversible \[51\]. This can be considered as an evidence for the lack of any electronic communication between the two ferrocenyl moieties.

However, for 1,2-diferrocenylethene and diferrocenylylene the second oxidation peak appears again and the separation potential values are of 160 and 140 mV, respectively. Thus the conjugated unsaturated C=C and C=C bonds allow again for the electronic communication between these two moieties \[51,53\].

As it is shown in Figure 5, thioketone 1a in dichloromethane solution, similarly to biferrocene \[48\], undergoes two distinct one-electron oxidations. Both signals correspond to the oxidation of the two iron atoms in the ferrocene moieties. First oxidation occurs at \(E_{\text{ox1}} = 0.164 \) V and the corresponding reduction is observed at \(E_{\text{red1}} = 0.099 \) V. The second oxidation occurs at \(E_{\text{ox2}} = 0.392 \) V with the corresponding reduction peak at \(E_{\text{red2}} = 0.332 \) V. Both red-ox systems are reversible because of the peak-to-peak separation which is \(\Delta E_1 = 65 \) mV and \(\Delta E_2 = 60 \) mV, respectively, as well as the \(I_{\text{ox}}/I_{\text{red}}\) ratios close to 1.

![Figure 5. Cyclic voltammetry of complex 1a (1.0 mmol/L) in CH\(_2\)Cl\(_2\)-nBu\(_4\)N][BF\(_4\)] (0.1 m/L) solutions at various scan rates (left) and at \(v = 0.75 \) V/s (right). The potentials \(E\) are given in V and referenced to the Fe\(^{III}\)/Fe\(^{II}\) couple. The arrows indicate the scan direction.](image-url)

The fact that the peak-to-peak separation \(\Delta E\) values slightly departed from the theoretical value of 59 mV, seem to contrast with diagnostic criteria for an electrochemically reversible one-electron
electrode process. Not completely compensated resistance given by the applied solvent and supporting solution can explain this fact.

The dependence of scan rate measurements showed that the peak currents for each peak varied linearly with the square root of scan rate ($v^{1/2}$) (Figure 6). It can point to the diffusion controlled process. Also, the dependence of logarithm of peak currents versus logarithm of scan rates displayed straight lines with slope values 0.5319 and 0.5767, respectively, that were close to the theoretical value of 0.5 which is expected for a diffusion controlled electrode process.

\[
K_{\text{comp}} = 10^{(F \times \Delta E/2.303 \times RT)}
\]

at 25 °C $K_{\text{comp}} = 10^{(38.92 \times \Delta E/2.303)} = 10^{(16.90 \times \Delta E)}$

for $\Delta E_0^0 = 0.392(V_{\text{ox2}}) \ - \ 0.164(V_{\text{ox1}}) \ V = 0.228 \ V$

$K_{\text{comp}} = 10^{(16.90 \times \Delta E)} = 10^{(16.90 \times 0.228)} = 10^{3.8532}$

According to Robin-Day classification [54], if $K_{\text{comp}}$ value is within the range $10^2$–$10^6$ the compound belongs to class II in which the charge is slightly delocalized between the two red-ox centers. An oxidation of the thiocarbonyl functionality was not observed within applied potential range. Thus, in contrast to ketones, the ferrocenyl moiety seems to stabilize the thiocarbonyl functionality in the molecule of ferrocenyl functionalized thioketones [53].
3.2.2. Platinathiiranes 3a–d

Complex 3d (Figure 7), bearing the ferrocene moiety in the bridging bisphosphane ligand shows an additional oxidation process at 0.757 V and reduction at 0.641 V ($\Delta E_p = 116$ mV). This irreversible process can be assigned to the dppf ligand. To prove this we also measure the electrochemical behavior of 2c which shows a reversible oxidation event at 0.651 V. It can be seen as the reversible oxidation of the ferrocene function in the dppf ligand chelating the platinum center. Its difference to the potential of the reversible oxidation of 3d is caused by the different $\pi$-backbonding of the norbornene to the (bisphosphane)Pt complex fragment.

![Figure 7](image_url)

**Figure 7.** Cyclic voltammetry of complex 3d (1.0 mmol/L) in CH$_2$Cl$_2$-[nBu$_4$N][BF$_4$] (0.1 m/L) solutions at various scan rates (left) and at $v = 0.50$ V/s (right). The potentials $E$ are given in V and referenced to the Fc$^+/Fc$ couple. The arrows indicate the scan direction.

The electrochemical properties of complexes 3a–c show a much more reversible character. The voltammograms of complex 3a depicted in Figure 8 show three full reversible oxidation events at 0.003 V, 0.363 V, and 0.668 V. The reversible oxidation of two ferrocene moieties at the thioketone site can be observed at 0.363 V and 0.668 V, respectively. The ferrocene moieties at the thioketone site seem to stabilize the oxidized intermediates of the sulfur and a chemical reaction leading to thiosulfine as oxidation products are hindered.

![Figure 8](image_url)

**Figure 8.** Cyclic voltammetry of complex 3a (1.0 mmol/L) in CH$_2$Cl$_2$-[nBu$_4$N][BF$_4$] (0.1 m/L) solutions at various scan rates (left) and at $v = 1.50$ V/s (right). The potentials $E$ are given in V and referenced to the Fc$^+/Fc$ couple. The arrows indicate the scan direction.

On the voltammograms of complex 3b presented in Figure 9 three reversible redox systems can be observed. The first well-defined oxidation at 0.390 V is followed by two poorly separated signals at 0.650 V and 0.750 V, respectively. In comparison to other compounds these signals correspond to the oxidation of the two ferrocene moieties of the thioketone ligand. Remarkably, oxidation of the sulfur atom is not observed in this complex.
The root-mean-square deviation (RMSD) of the calculated bond lengths from the corresponding experimental data amounts to 0.065 Å. Even smaller deviation can be reached for the bond lengths involving only non-hydrogen atoms. Similarly, the calculated bond angles differ from the corresponding experimental values by no more than 0.92° on average (RMSD = 1.01° for the bond angles involving only non-hydrogen atoms). The torsion angles of the optimized complex show however significant deviations from their XRD values (RMSD = 15.0° and 18.1° for the torsion angles involving all and only non-hydrogen atoms, respectively). The greatest differences occur for the torsion angles defining the spatial arrangement of all phenyl rings and two ferrocenyl groups at the C(1) atom. These groups are located further away from the rigid core of 3c and their torsion angles are softer. In consequence,
they are most affected by crystal packing effects, as it is manifested by the XRD structure. The optimized geometry of the complex 3c displays an almost planar fragment containing the Pt center and its four adjacent atoms [that is, P(1), P(2), S(1), and C(1)]. The calculated bond lengths and bond angles in the immediate neighborhood of the Pt center are listed in Table 1. The calculated bond lengths and angles involving the Pt center differ by no more than 0.032 Å and 3.4° from the corresponding experimental results. The Pt-P(1) and Pt-P(2) bonds are elongated slightly upon complexation of (dppf)Pt with 1a. Similarly, the C=S bond of 1a becomes longer if coordinated by the Pt center of (dppf)Pt.

Table 1. Calculated coordination geometry around the Pt center in complex 3c. For some of the geometrical parameters, their values calculated for the isolated 1a and (dppf)Pt molecules are presented in parentheses.

<table>
<thead>
<tr>
<th>Atoms</th>
<th>Bond Length (Å)</th>
<th>Atoms</th>
<th>Bond Angle (°)</th>
</tr>
</thead>
<tbody>
<tr>
<td>Pt-P(1)</td>
<td>2.246 (2.213)</td>
<td>P(1)-Pt-P(2)</td>
<td>99.8 (129.1)</td>
</tr>
<tr>
<td>Pt-P(2)</td>
<td>2.282 (2.213)</td>
<td>P(1)-Pt-C(1)</td>
<td>105.7</td>
</tr>
<tr>
<td>Pt-S(1)</td>
<td>2.326</td>
<td>P(2)-Pt-S(1)</td>
<td>107.7</td>
</tr>
<tr>
<td>Pt-C(1)</td>
<td>2.155</td>
<td>C(1)-Pt-S(1)</td>
<td>46.8</td>
</tr>
<tr>
<td>C(1)-S(1)</td>
<td>1.788 (1.665)</td>
<td></td>
<td></td>
</tr>
</tbody>
</table>

Table 2 presents the Wiberg bond index (WBI) [56] and the delocalization index (DI) [57], which are obtained from the NBO and QTAIM calculations, respectively. The elongations of the Pt-P(1), Pt-P(2), and C(1)-S(1) bonds in 3c are associated with decreases in their bond order indices. Both WBI and DI suggest that the Pt-S(1) bond is essentially similar to Pt-P(1) and Pt-P(2). Compared to these bonds, the Pt-C(1) bond is characterized by lower values of WBI and DI. This implies smaller strength and covalent character for the Pt-C(1) bond. The formation of the complex results in particularly significant decreases in the WBI and DI values of the C(1)-S(1) bond. These indices reveal that the C(1)-S(1) bond order loses its formally double character and this bond actually becomes close to a single bond (a typical single C-S bond, such as that in tert-butylthiol, shows the WBI and DI values of 0.974 and 1.023, respectively).

Table 2. Calculated natural bond orbitals (NBO) and quantum theory of atoms in molecules (QTAIM) indicators of bond order for the bonds around the Pt center in complex 3c. For some of the bonds, their order indices calculated for the isolated 1a and (dppf)Pt are presented in parentheses.

<table>
<thead>
<tr>
<th>Atoms</th>
<th>WBI</th>
<th>DI</th>
</tr>
</thead>
<tbody>
<tr>
<td>Pt-P(1)</td>
<td>0.654 (0.835)</td>
<td>1.057 (1.1865)</td>
</tr>
<tr>
<td>Pt-P(2)</td>
<td>0.636 (0.835)</td>
<td>0.999 (1.1869)</td>
</tr>
<tr>
<td>Pt-S(1)</td>
<td>0.689</td>
<td>0.926</td>
</tr>
<tr>
<td>Pt-C(1)</td>
<td>0.417</td>
<td>0.701</td>
</tr>
<tr>
<td>C(1)-S(1)</td>
<td>1.148 (1.656)</td>
<td>1.157 (1.696)</td>
</tr>
</tbody>
</table>

The bonds around the Pt center are further characterized using QTAIM parameters determined at bond critical point (BCP). Bond paths and associated BCPs were found for all four bonds involving the Pt center. The values of several important QTAIM parameters at the BCPs are listed in Table 3.
In general, the QTAIM characteristics of the Pt-P(1), Pt-P(2), Pt-S(1), and Pt-C(1) bonds are rather typical of metal-ligand bonds [58–60]. According to QTAIM, metal-ligand bonds are classified as interactions falling in the transit zone between shared (that is, classical covalent) and purely closed shell (that is, ionic) interactions [61]. On the one hand, the BCPs of Pt-P(1), Pt-P(2), Pt-S(1), and Pt-C(1) show relatively small values of the electron density ($\rho$) and the positive sign of the Laplacian of $\rho$ ($\nabla^2 \rho$) [61]. Such features reveal the closed shell character of interactions. On the other hand, the total energy density ($\rho$) at these BCPs adopts negative values, which in turn implies a covalent contribution to bonding in Pt-P(1), Pt-P(2), Pt-S(1), and Pt-C(1) [62]. The presence of covalent contribution is confirmed by two other BCP parameters: the ratio of the potential energy density ($V$) to the kinetic energy density ($G$) and the ratio of two Hessian eigen values ($\lambda_1/\lambda_3$). For the four bonds, their values of $-V/G$ are between 1 and 2 and $-\lambda_1/\lambda_3 > 0.25$. Such ranges of $-V/G$ and $-\lambda_1/\lambda_3$ values indicate some degree of covalency between bonded atoms [63,64]. The BCP of the C(1)-S(1) bond in both the complex and 1a possesses the features of typical shared interactions (a high value of $\rho$, $\nabla^2 \rho < 0$, $H < 0$). Furthermore, its QTAIM parameters demonstrate a decrease in electron sharing between C and S upon complexation with (dppf)Pt.

Contours of frontier molecular orbitals for 3c are depicted in Figure 11. The main part of the highest occupied molecular orbital (HOMO) is localized around the Fe atoms of thioketone fragment. The contribution of Pt orbitals to the HOMO is also evident. In general, both the HOMO and its nearest orbitals with lower energies (e.g., HOMO-1, HOMO-2, HOMO-3, see Figure S3) are largely composed of the Fe orbitals of thioketone fragment.

Table 3. Calculated QTAIM parameters of bond critical points for the bonds around the Pt center in complex 3c. For some of the bonds, their QTAIM parameters calculated for the isolated 1a and (dppf)Pt are presented in parentheses.

<table>
<thead>
<tr>
<th>Atoms</th>
<th>$\rho$ (au)</th>
<th>$\nabla^2 \rho$ (au)</th>
<th>$H$ (au)</th>
<th>$-V/G$</th>
<th>$-\lambda_1/\lambda_3$</th>
</tr>
</thead>
<tbody>
<tr>
<td>Pt-P(1)</td>
<td>0.1232</td>
<td>0.0755</td>
<td>-0.0642</td>
<td>1.7733</td>
<td>0.3711</td>
</tr>
<tr>
<td></td>
<td>(0.1278)</td>
<td>(0.0584)</td>
<td>(-0.0712)</td>
<td>(1.8305)</td>
<td>(0.4011)</td>
</tr>
<tr>
<td>Pt-P(2)</td>
<td>0.1151</td>
<td>0.0911</td>
<td>-0.0557</td>
<td>1.7104</td>
<td>0.3413</td>
</tr>
<tr>
<td></td>
<td>(0.1279)</td>
<td>(0.0584)</td>
<td>(-0.0713)</td>
<td>(1.8306)</td>
<td>(0.4011)</td>
</tr>
<tr>
<td>Pt-S(1)</td>
<td>0.1010</td>
<td>0.1457</td>
<td>-0.0401</td>
<td>1.5242</td>
<td>0.2912</td>
</tr>
<tr>
<td>Pt-C(1)</td>
<td>0.1100</td>
<td>0.1239</td>
<td>-0.0454</td>
<td>1.5951</td>
<td>0.3625</td>
</tr>
<tr>
<td>C(1)-S(1)</td>
<td>0.1839</td>
<td>-0.2656</td>
<td>-0.1349</td>
<td>2.9701</td>
<td>1.1881</td>
</tr>
<tr>
<td></td>
<td>(0.2269)</td>
<td>(-0.3901)</td>
<td>(-0.2612)</td>
<td>(2.5958)</td>
<td>(2.1543)</td>
</tr>
</tbody>
</table>

Figure 11. Contours of the highest occupied molecular orbital and lowest unoccupied molecular orbital for complex 3c. Their parts possessing either positive or negative sign are colored red and green, respectively. The contours are plotted with an isovalue of 0.05 au.
In contrast to the HOMO, the lowest unoccupied molecular orbital (LUMO) is mainly concentrated on the ferrocenyl and phenyl groups of (dppf)Pt fragment. Similarly, the molecular orbitals above the LUMO also show significant contributions from these groups (see Figure S4). The complex exhibits a HOMO-LUMO gap of 1.82 eV.

In addition to the visual inspection of HOMO and LUMO contours, the frontier molecular orbitals of 3c are also analyzed quantitatively by expressing their percentage composition in terms of natural atomic orbitals (NAOs). The HOMO mainly consists of the valence 3d NAOs belonging to two Fe atoms of thioketone fragment. The percentage contributions from the NAOs of these Fe atoms sum up to ca. 73% in the HOMO. The leading contribution of ca. 29% originates from the 3d\(_{yz}\) orbital of one of the Fe atoms. Next relevant contributions, but much smaller in magnitude, are ascribed to the NAOs of the Pt and S(1) atoms. These atoms make similar contributions, each adds up to ca. 3%. The LUMO shows a significant contribution of the 3d\(_{x^2−y^2}\) NAO belonging to the Fe atom of (dppf)Pt fragment. This NAO constitutes ca. 21% of the LUMO and it is by far the largest individual NAO contribution. In total, the Fe atom of (dppf)Pt makes a contribution of ca. 32%. The rest of the LUMO also comes from the (dppf)Pt fragment, in particular from the NAOs of C atoms belonging to its ferrocenyl and phenyl groups.

The NBO analysis carried out for 3c reveals that there is only one bonding NBO in the thiocarbonyl group of thioketone fragment. Within the framework of the NBO method, bonding and back-bonding contributions to metal-ligand bonds can be described by electron delocalizations between donor and acceptor NBOs (such delocalizations can be further characterized by the respective stabilization energies derived from the second-order perturbation theory of NBO donor-acceptor interactions [38]). Donor-acceptor delocalizations from the bonding NBO of C(1)-S(1) to the lone pair and Rydberg NBOs of the Pt center have indeed been detected in 3c but their stabilization energies are not dominating (ca. 55 kcal/mol in total). Very small stabilization energies (ca. 9 kcal/mol in total) are attributed to the delocalizations from the valence NBOs of Pt to the antibonding NBO of C(1)-S(1). In this case, the valence NBOs of Pt include its 5d\(_{z^2}\) and 5d\(_{x^2−y^2}\) orbitals.

It is instructive to relate the above computations to the results of CV measurements for 3c. As it was mentioned above, the HOMO of the complex is mainly localized on two Fe atoms of thioketone fragment, but their contributions to the HOMO are by no means equal. This proves the CV assignment of two lowest oxidation events to two ferrocene moieties at the thioketone site of 3c. For various diferrocenyl derivatives it was reported that the separation between their oxidation potentials is often associated with the distance between their Fe atoms [35]. The separation tends to increase as the Fe–Fe distance decreases. For thioketone 1a, its Fe–Fe distance and oxidation potential separation amount to 5.207 Å and 0.228 V, respectively. For complex 3c in its optimized geometry, the distance between two Fe atoms of the thioketone fragment is 5.125 Å and the value of 0.306 V was measured as the separation between two lowest oxidation potentials. Accordingly, the Fe–Fe distance is smaller for 3c, and therefore, the separation between its two lowest oxidation potentials grows, compared to 1a.

4. Conclusions

The present study showed that diferrocenyl thioketone (1a) reacts easily with (bisphosphane)Pt(0) complexes 2a–c, yielding platinathiiranes 3a–c. In contrast to complexes 3a,b, complex 3c possesses an additional ferrocenyl group which acts as the backbone of the bisphosphane ligand. In addition, platinathiirane 3d was prepared from thiobenzophenone (1b) and 2c. Three of the four complexes 3a–d have been characterized using XRD methods revealing that the bond lengths in the platinathiirane rings correspond to the data reported for similar platinaheterocycles derived from several aryl and cycloaliphatic thioketones [16–18,47]. Compounds 3a–d were characterized electrochemically using CV. For the sake of comparison, the starting thioketone 1a, was also studied electrochemically. Complexes 3a–c derived from 1a behaved similarly and exhibit reversible oxidation processes. Together with 3d, these complexes undergo a mixed diffusion/adsorption controlled electrode process, whereas 1a is oxidized by a diffusion controlled mechanism. The quantum-chemical computations performed for 3c
revealed that the C=S bond of \( \text{1a} \) lost its formally double character upon complexation with (dppf)Pt and its bond order became quite close to unity.

The chemistry of metallathiiranes and related metalla-thiirenes attracts considerable attention of many groups and the nature of the metal-C=S bond in these compounds is a problem of current interest [65]. In addition, because of the well-documented importance of ferrocenyl functionalized compounds, the new complexes \( \text{3a-d} \) offer a novel type of ferrocene- and platinum-containing sulfur systems for potential applications in material chemistry [66]. Notably, some of the Pt sulfur/nitrogen complexes were reported to induce stereoselective polymerization in methylthiirane [67] but in general, the problem of potential application of sulfur–platinum complexes in polymer chemistry is lesser known and requires further studies.

In the final conclusion it should be also stressed that the present study evidences one more time the importance of thioketones for organometallic and coordination chemistry. In recent two decades they were demonstrated to act as perfect complexation reagents for diverse metals at variable oxidation states, such as platinum, ruthenium, tungsten, etc., [68]. For that reason, ferrocenyl functionalized thioketones reported in the study should be also regarded as potentially useful trapping reagents not only for Pt(0) complexes but also for other metals. Potentially, they can also be applied in reactions with iron carbonyls aimed at the preparation of new redox-active iron–sulfur clusters recognized as hydrogenase mimics [69–72].

**Supplementary Materials:** The following are available online at http://www.mdpi.com/1996-1944/12/17/2832/s1, Section S1: Selected bond lengths and angles measured for complexes \( \text{3a-c} \) using XRD methods (Table S1) and crystal data and refinement details for the X-ray structure determinations (Table S2). Section S2: Further electrochemical details: The dependence of the peak currents \( (I) \) from the square root of scan rate \( (\sqrt{v}) \) and logarithm of peak currents \( (\log I) \) versus logarithm of scan rates \( (\log v) \) for complexes \( \text{3a-d} \) (1.0 mmol/L) in \( \text{CH}_2\text{Cl}_2-\text{(nBu}_4\text{N})\text{[BF}_4] \) (0.1 m/L) (Figure S1). Section S2: Further computational details. Table S3: Calculated NBO and QTAIM indicators of bond order for the bonds around the Pt center in complex \( \text{3c} \). Table S4: Calculated QTAIM parameters of bond critical points for the bonds around the Pt center in complex \( \text{3c} \). Figure S2: Superposition of the calculated and experimental structures of complex \( \text{3c} \). Figure S3: Contours of several molecular orbitals below and above the frontier molecular orbitals for complex \( \text{3c} \). Figure S4: Contours of the highest occupied molecular orbital and lowest unoccupied molecular orbital for complex \( \text{3c} \). Section S4: Cartesian coordinates (in Å) for \( \text{3c} \), \( \text{1a} \), and (dppf)Pt-fragment: Tables S5–S7.


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**Conflicts of Interest:** The authors declare no conflict of interest.

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