

Short Note

Di(bicyclo[4.2.0]octa-1(6),2,4-trien-3-yl)dimethylsilane

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Abstract: Di(bicyclo[4.2.0]octa-1(6),2,4-trien-3-yl)dimethylsilane (**DiBCB-DMS**) was synthesized as a promising monomer for polymer coating production with good dielectric properties for electronics. **DiBCB-DMS** was characterized by ¹H, ¹³C NMR, IR, UV/VIS, and HRMS analysis.

Keywords: di(bicyclo[4.2.0]octa-1(6),2,4-trien-3-yl)dimethylsilane; benzocyclobutene (**BCB**); monomer; polymers; dielectric coating

1. Introduction

The miniaturization of electronic devices encourages the creation of multi-level integrated circuits with high functional density. It is possible due to the use of low-dielectric-constant (low-k) materials incorporated into interconnected media. Traditional dielectric materials based on inorganic silica are inferior to organic polymer materials due to their significant advantages in solution processability and mechanical flexibility [1–4].

Organosilicon resins with benzocyclobutene (**BCB**) filler have been demonstrated to be promising as low-dielectric-constant (low-k) materials for the electronics industry and related applications. These materials are currently being actively studied [1–6].

In particular, Dow Chemicals (www.dow.com) presents a line of composite commercial materials (CYCLOTENE™) based on divinyltetramethylsiloxane-benzocyclobutene (**DVS-BCB**) as photoresists and thermosets.

Recently, a direction has been developed associated with the use of multi-benzocyclobutene functionalized silane and siloxane monomers, which provide more dense crosslinking, high thermal stability, and low coefficient of thermal expansion. In particular, polymers are known based on tri-(benzocyclobuten-4-yl) phenylsilane and tetra-(benzocyclobuten-4-yl)silane [7] or their siloxane analogs [6]. However, it is a surprise that there are no data about the synthesis or physical and chemical properties of a simpler analogue, namely di(bicyclo[4.2.0]octa-1(6),2,4-trien-3-yl)dimethylsilane.

2. Results

In this paper, we present a simple preparation procedure and the physical properties of a thermopolymerizable monomer containing two benzocyclobutene and dimethylsilane fragments (di(bicyclo[4.2.0]octa-1(6),2,4-trien-3-yl)dimethylsilane). The synthesis involves a two-stage process starting from 4-bromobenzocyclobutene obtained by a previously described method [8,9]. In the first stage, 4-bromobenzocyclobutene (**1**) reacts with magnesium in the presence of iodine in anhydrous

tetrahydrofuran to form a Grignard reagent (2) (Figure 1). Further addition of dichlorodimethylsilane under ice-cooling (0–10 °C) leads to the formation of the target product **diBCB-DMS** in 70% yield. The product was purified by distillation under reduced pressure (150–160 °C at 0.5–0.7 mbar) and obtained as a transparent colorless liquid crystallizing in the refrigerator.

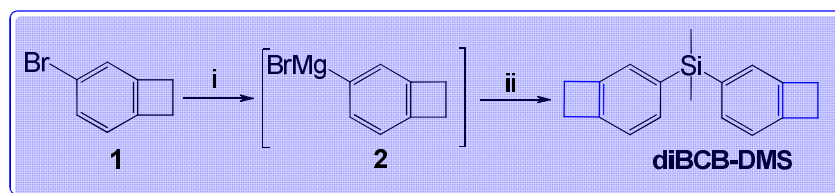


Figure 1. Scheme of the synthesis of **diBCB-DMS**: (i) Mg, I₂, THF, 25–45 °C, 2 h; (ii) (Me)₂SiCl₂, 0–5 °C, 1 h, r.t. 12 h.

DiBCB-DMS was characterized by ¹H, ¹³C NMR, IR, UV/VIS, and HRMS analysis. These data can be found in the Supplementary Materials.

The resulting monomer **diBCB-DMS** is polymerized when heated above 200 °C. It is known that the mechanism of polymerization of benzocyclobutene derivatives is multivariable [2,4,6,7,10]. Benzocyclobutene fragments in thermo-induced o-xylylene form can react with active double bonds. Thermopolymerization in this case proceeds according to the [4+2] cycloaddition mechanism. In the absence of compounds with active double bonds in the reaction mass, the benzocyclobutene ring opens and reacts with another similar fragment in different ways. The scheme of the estimated polymerization processes is shown in Figure 2.

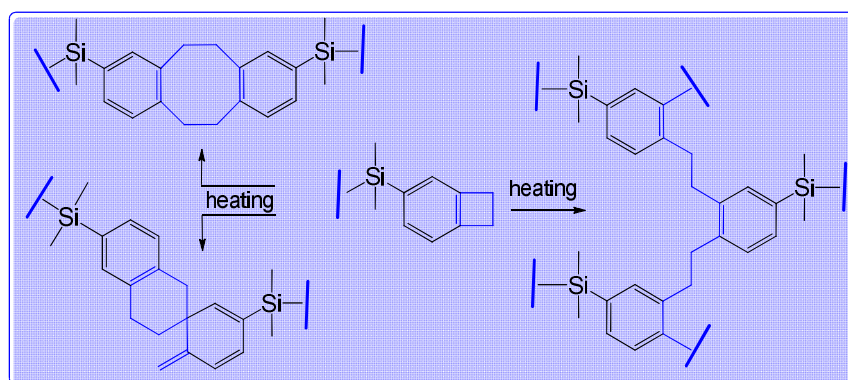


Figure 2. Scheme of the proposed polymerization mechanism of **diBCB-DMS**.

The colorless liquid monomer **diBCB-DMS** gave a brown solid polymer upon heating at 250 °C for 30 min in air. Color change and characteristic adsorption bands of the polymer indicated that remarkable oxidative processes occurred. A decrease in mass was also observed. Therefore, a step-by-step polymerization process (160 °C–1 h, 180 °C–2 h, 200 °C–2 h, 220 °C–2 h, and 240 °C–30 min) under Ar was applied, giving a white to light yellow polymer without decrease in mass.

The resulting polymers were studied by IR spectroscopy. IR spectra of the monomer and polymers are presented in the Supplementary Materials. According to the literature [11,12], the band at 1466 cm⁻¹ characterizes the vibrations of the CH₂ groups of the cyclobutene ring in monomer molecules and the band at 1493 cm⁻¹ characterizes the vibrations of dibenzocyclooctadiene and tetrahydronaphthalene structures in the crosslinked polymer. Moreover, characteristic adsorption bands of carbonyl at about 1600–1776 cm⁻¹ and OH groups at >3500 cm⁻¹ occurred, indicating the oxidation of cyclobutene moieties. The oxidation may occur primarily at the tetrahydronaphthalene group formed during the Diels–Alder reaction of BCB [13].

We assume that the polymerization of **diBCB-DMS** proceeds, giving a crosslinked polymer. This fact is indirectly confirmed by high resistance of the obtained polymer to the action of organic solvents and mechanical strength at high temperatures.

3. Materials and Methods

All reagents were purchased from Sigma-Aldrich (St. Louis, MO, USA) (Lancaster, UK) and used without further purification unless otherwise stated. Dimethyldichlorosilane was purchased from abcr GmbH (Karlsruhe, Germany).

NMR spectra were registered using a Bruker AM-300 spectrometer (Bruker Corporation, Billerica, MA, USA) in CDCl₃. Mass spectra were obtained on a Varian MAT CH-6 instrument (Varian, Inc., Palo Alto, CA, USA) using a direct inlet system; the ionization energy was 70 eV and the acceleration voltage was 1.75 kV. The reaction mixtures were analyzed and the purity of all products was checked by TLC on Merck Silica gel 60 F254 UV-254 plates (Darmstadt, Germany). The optical properties of **diBCB-DMS** were studied in acetonitrile solution (Sigma-Aldrich, St. Louis, MO, USA) using a Lambda 35 UV/VIS spectrometer (PerkinElmer, Singapore). Absorption spectra were measured at a concentration of 0.015–0.003 mg/mL and an optical path length of 0.2 cm.

4. Synthesis of diBCB-DMS

4-Bromobenzocyclobutene (18.3 g, 0.1 mol) was added dropwise with vigorous stirring to magnesium powder (2.4 g, 0.1 mol), activated by the addition of an iodine crystal in 100 mL of freshly distilled dry tetrahydrofuran. After the start of the reaction, 4-bromobenzocyclobutene was added at such a rate that the temperature of the reaction mixture was 40–45 °C. When the addition was complete, the reaction mass was stirred for 2 h at this temperature and then was cooled to 0 °C. Dichlorodimethylsilane (6.4 g, 0.1 mol) was added dropwise at 0–10 °C and the mixture was allowed to warm to room temperature. The next day, the mixture was diluted with hexane (100 mL). The precipitate was filtered off and, washed with several portions of hexane. The filtrate was concentrated on a rotary evaporator. The remaining crude product was purified by vacuum distillation at a distillation temperature of 150–160 °C (0.5–0.7 mbar). The yield was 9.2 g (70%). **diBCB-DMS** is a colorless liquid solidifying in the refrigerator. ¹H NMR (300 MHz, CDCl₃) δ 7.62 (d, *J* = 7.3 Hz, 2H), 7.46 (s, 2H), 7.28 (d, *J* = 7.3 Hz, 2H), 3.40 (s, *J* = 3.4 Hz, 8H), 0.76 (s, 6H). ¹³C NMR (76 MHz, CDCl₃) δ 147.30, 145.75, 136.85, 132.73, 128.16, 122.13, 30.14, 30.01, -1.67. HRMS (+MS) found: 265.1414; calculated: 265.1407.

5. Conclusions

New di(bicyclo[4.2.0]octa-1(6),2,4-trien-3-yl)dimethylsilane (**DiBCB-DMS**) was obtained via a two-stage synthesis starting from 4-bromobenzocyclobutene with good yields. **DiBCB-DMS** was characterized by ¹H, ¹³C NMR, IR, UV/VIS, and HRMS analysis. The obtained product can be useful for the development of new polymerizable composite materials with good thermal stability and dielectric properties.

Supplementary Materials: Supplementary data to this article can be found online. These data include MOL file, ¹H, ¹³C NMR, HRMS, UV/VIS, and IR of the title compound.

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Conflicts of Interest: The authors declare no conflict of interest. The funders had no role in the design of the study; in the collection, analyses, or interpretation of data; in the writing of the manuscript; or in the decision to publish the results.

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